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## Calorimetry Lab

### Calculations

Lab 2. Coffee Cup

Calorimetry

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Calorimeters can be used to find a substance's specific heat capacity. You will use the Calorimetry Lab Gizmo™ to determine the specific heat capacities of various substances. 1. On the SIMULATION pane,

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select Copper.

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Correct Answer: B 1 is a pain receptor, 2 is a temperature receptor, 3 is a light touch receptor, and 4 is a strong pressure receptor

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Investigate how calorimetry can be used to find relative specific heat values when different substances are mixed with water.

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Vocabulary: calorie,

calorimeter, joule,

specific heat capacity

Prior Knowledge

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Gizmo Questions (Do these  
BEFORE using the  
Gizmo.)

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calorimeter and heated  
2,000 g of water by  $20^{\circ}$   
C. How much thermal  
energy has been  
released? (Hint: Use a  
specific heat equation.)  
Give your answer in  
both joules and calories.

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C. How many kilocalories does the snack bar contain? In this Gizmo, you will investigate how calorimetry can be used to find relatively

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calorimeter is an



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insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up.

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Investigate how

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Calorimetry can be used to find relative specific heat values when different substances are mixed with water.

Modify initial mass and temperature values to see effects on the system. One or any combination of the substances can be mixed with water. A dynamic graph (temperature vs. time) shows

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temperatures of the individual substances after mixing.

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Calorimeters can be used to find a substance's specific heat capacity. You will use the Calorimetry Lab Gizmo™ to determine the specific

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heat capacities of various substances. 1.

On the SIMULATION pane, select Copper.

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Calorimeters can be used to find a substance's specific heat capacity. You will use the Calorimetry Lab Gizmo™ to

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determine the specific heat capacities of various substances. 1. On the SIMULATION pane, select Copper.

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Correct Answer: C. The final temperature of the lead-water system will be lower than the final temperature of the

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## Calorimetry Lab

### Copper-water system. A

blacksmith heats a 1,540 g iron horseshoe to a temperature of  $1445^{\circ}\text{C}$  before dropping it into 4,280 g of water at  $23.1^{\circ}\text{C}$ .

### ~~Calorimetry Lab~~ Flashcards | Quizlet

Investigate how calorimetry can be used to find relative specific heat values when

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different substances are mixed with water.

Modify initial mass and temperature values to see effects on the system. One or any combination of the substances can be mixed with water. A dynamic graph (temperature vs. time) shows temperatures of the individual substances after mixing.

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## Calorimetry Lab

### Gizmo Warm-up Answers

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up. Calorimeters can be used to find a substance's specific

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~~+ Calorie + Heat~~

Gizmo Warm-up A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water

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