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is by performing a titration on a sample of blood. In this process, the alcohol from the blood is oxidized by dichromate ions ($\text{Cr}_2\text{O}_7^{2-}$) according to the following net ionic equation: $\text{C}_2\text{H}_5\text{OH} + 2\text{Cr}_2\text{O}_7^{2-} + 16\text{H}^+ \rightarrow 2\text{CO}_2 + 4\text{Cr}^{3+} + 11\text{H}_2\text{O}$ A 10.00-g sample of blood was drawn from a patient, and 13.77 mL of 0.02538 M $\text{K}_2\text{Cr}_2\text{O}_7$ was required to titrate the alcohol.

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